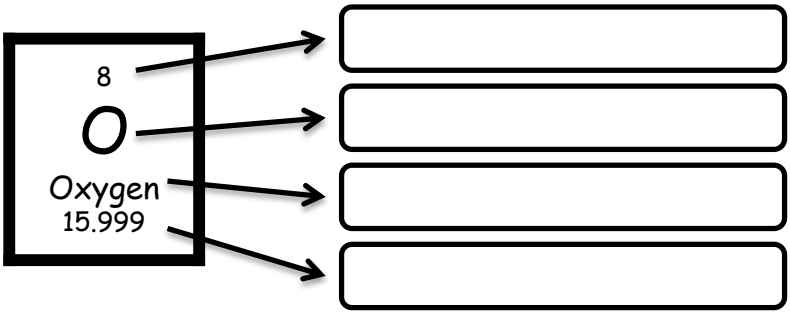
Name: _____ **Date:** _____ **B#:** ____

Atomic Structure:



What 2 parts of an atom does the atomic # represent?

_____ & _____

How do you figure out the # of neutrons?

_____

6 C Carbon 12.011	# P: ____ # E: ____ # N: ____	10 Ne Neon 20.179	# P: ____ # E: ____ # N: ____	19 K Potassium 39.098	# P: ____ # E: ____ # N: ____

CREATING BOHR DIAGRAMS

Rules for arranging electrons:

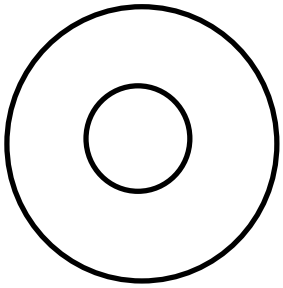
1. The 1st energy level can hold up to 2 electrons.
2. The 2nd energy level can hold up to 8 electrons.
3. The 3rd energy level can hold up to 8 electrons.

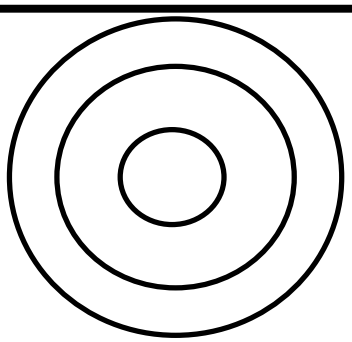
What term is used to describe the electrons in the outermost energy level?

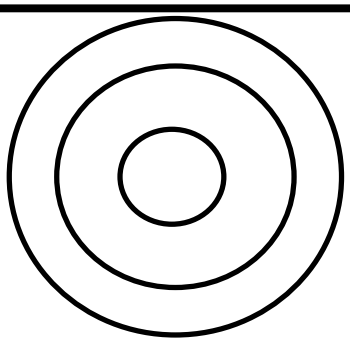
_____

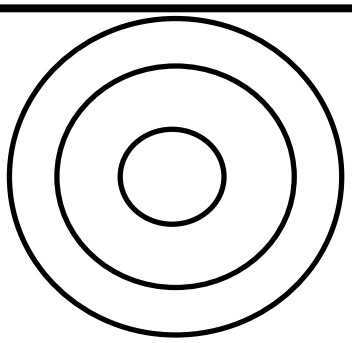
Sketch An Atom	
Draw 5 protons in the nucleus and label with the charge.	
Draw 6 neutrons in the nucleus and label with the charge.	
Draw 2 electrons in the 1st energy level and label with their charge.	
Draw 3 electrons in the 2nd energy level and label with their charge.	
What element is represented?	

Sketch An Atom	
Draw 3 protons in the nucleus and label with the charge.	
Draw 4 neutrons in the nucleus and label with the charge.	
Draw 2 electrons in the 1st energy level and label with their charge.	
Draw 1 electrons in the 2nd energy level and label with their charge.	
What element is represented?	

Neon		
# P		
# E		
# N		
# of Valence Electrons		
Atomic #: 10		Mass #: 20

Magnesium		
# P		
# E		
# N		
# of Valence Electrons		
Atomic #: 12		Mass #: 24

Chlorine		
# P		
# E		
# N		
# of Valence Electrons		
Atomic #: 17		Mass #: 35

Silicon		
# P		
# E		
# N		
# of Valence Electrons		
Atomic #: 14		Mass #: 28

CREATING LEWIS DOT DIAGRAMS

Rules for arranging electrons:

1. Figure out how many valence electrons the element has in its atom.
2. Place dots around the element's symbol one at a time (can't exceed 8).

Ne **Mg** **Cl** **Si**