

HW10, chemistry 2

Factors affecting reaction rate:

Concentrations

Pressure (for reactions involving gases)

Surface area of solid reactants

Temperature

Catalysis

Temperature: as the temperature increases, the rate of reaction increases following this equation (every increase by 10 degrees in temperature will increase reaction rate (or rate constant) by 2 to 4 times):

$$k_{t+10} / k_t = \gamma$$

$$\gamma = 2-4$$

t – temperature in C (or K)

Catalysis:

A catalyst is a substance that increases the rate of chemical reaction without being used up, a catalyst acts by allowing the reaction to proceed by alternative pathway with lower activation energy. Metals, oxides, acids can be catalysts, as well as enzymes in biochemical processes.

Questions:

1. By how much will the rate constant increase if we raise the temperature by 30 degrees, given that γ is equal to 4?

2. $\text{Mg (s)} + 2\text{HCl (aq)} \rightarrow \text{MgCl}_2 + \text{H}_2$

Which of the following will not increase the rate of the reaction:

- Increasing the temperature
- Increasing the surface area of Mg
- Increasing the volume of HCl
- Increasing the concentration of HCl

3. MgO₂ is a catalyst in the decomposition of hydrogen peroxide (H₂O₂). Which of the following best describes how a catalyst works?

- It speeds up the reaction by increasing the average energy of the particles
- It speeds up the reaction by allowing the reaction to occur by an alternative pathway of lower activation energy
- It speeds up the reaction by increasing the average energy
- It slows down the reaction by decreasing the collision frequency of particles